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Chapter 14 Acids
And Bases

Chapter 14 Acids And Bases

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an acid that is
composed of H, O, and
a third element. The H
& O are covalently
bonded and are called
the hydroxyl group.
Property of bases.
taste bitter, feel
slippery, change color
of indicators, react with
acids to produce a salt

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Chapter 14 Acids And Bases

and water

neutralization rxn, are electrolytes. Arrhenius or traditional base.

Chapter 14 Acids and Bases

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Chapter 14. Table of Contents. Copyright ©2017 Cengage Learning. All Rights Reserved. (14.1) The nature of acids and bases (14.2) Acid strength (14.3) The pH scale (14.4) Calculating

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the pH of strong acid solutions (14.5)
Calculating the pH of weak acid solutions (14.6) Bases (14.7)
Polyprotic acids.

Chapter 14 Acids and Bases - Accountax

CHAPTER 14 ACIDS
AND BASES 501 24. A
Lewis acid must have
an empty orbital to
accept an electron
pair, and a Lewis base
must have an unshared

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Chapter 14 Acids
And Bases
pair of electrons.

**CHAPTER 14 ACIDS
AND BASES -
learning.hccs.edu**

ANS: e) ClO_4^- , F^- ,
 OAc^- , CN^- PAGE: 14.2

11. The conjugate base
of a weak acid is a) a
strong base b) a weak
base c) a strong acid d)
a weak acid e) none of
these ANS: b) a weak
base PAGE: 14.2 12.

Which of the following
is the equilibrium
constant expression for

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the dissociation of the weak acid HOCl? a) $K = \frac{[H^+][OCl^-]}{[HOCl]}$

CHAPTER 14 Acids and Bases

Chapter 14 - Acids and Bases. Description.

Chapter 14 Review.

Total Cards. 36.

Subject. Chemistry.

Level. 11th Grade. ...

General Properties of Acids: Definition. 1)

sour taste, corrosive 2)

characteristically

changes acid-base

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indicator 3) react with active metals to form hydrogen gas 4) react with bases to form salt and water 5) conduct ...

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Acids and Bases

CHAPTER 14 BIG IDEA

Acids are substances that donate hydrogen ions in aqueous solutions. Bases are substances that accept

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hydrogen ions in
aqueous solutions.

SECTION 1 Properties
of Acids and Bases

SECTION 2 Acid-Base
Theories SECTION 3

Acid-Base Reactions

Why It Matters Video

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CorrectionKey=NL-A

DO NOT

**EDIT--Changes must
be made ...**

Chapter 14 - Acids and
Page 10/24

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Bases. STUDY.

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PLAY. Match. Gravity.

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ariannafletcher-Key

Concepts: Terms in this

set (41) T or F: In

solution, strong acids

and bases ionize

completely, but weak

acids and bases ionize

only partially. True. T

or F: An acidic solution

contains more

hydroxide ions than ...

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**Chapter 14 - Acids
and Bases**

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Modern Chemistry

Chapter 14 Vocab

Acids and Bases.

binary acid. oxyacid.

Arrhenius acid.

Arrhenius base. an acid

that contains only two

different elements:

hydrogen an.... an acid

that is a compound of

hydrogen, oxygen, and

a third el.... a chemical

compound that

increases the

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concentration of
hydrog....

chemistry chapter 14 acids bases Flashcards and Study Sets ...

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section 14.1 in
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14-Introduction ...

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Bases, & Salts
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Chapter 14 Acids And Bases

14.1 The Nature of Acids and Bases A. Arrhenius Model 1. Acids produce hydrogen ions in aqueous solutions 2. Bases produce hydroxide ions in aqueous solutions B. Bronsted-Lowry Model 1.

Chapter 14 - Acids and Bases - ScienceGeek.net

Chapter 14: Acids and
Bases Ch 14 Page 1
Page 15/24

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Chapter 14 Acids And Bases

Arrhenius Base - a hydroxide containing compound that dissociates to produce OH⁻ when dissolved in water $\text{NaOH(aq)} \rightarrow \text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq})$
Bases that contain OH⁻ are ionic compounds. Ionic substances dissociate in water.

Chapter 14: Acids and Bases

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Acids and Bases (Chapter 14)

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Chapter 14: Acids and Bases. Acids and Bases. Know the definition of Arrhenius, Bronsted-Lowry, and Lewis acid and base. Autoionization of Water. Since we will be dealing with aqueous acid and base solution,

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first we must examine the behavior of water.

1) In pure water, to which no electrolytes have been added, there is a small electrical conductivity, indicating that there are ions present.

Chapter 14: Acids and Bases

CHAPTER 14 REVIEW

Acids and Bases

SECTION 1 SHORT

ANSWER Answer the following questions in

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the space provided. 1.
Name the following
compounds as acids:
sulfuric acid a. H_2SO_4
sulfurous acid b. H_2SO_3
hydrosulfuric acid c.
 H_2S perchloric acid d.
 HClO_4 hydrocyanic
acid e. hydrogen
cyanide 2. H

14 Acids and Bases

Chapter Fourteen:

Acids and Bases

Chapter interactives:

Acids/bases and

indicators Chemistry

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battleship Chemistry

jokes pH simulation

Chapter Homework:

Section 1: Chapter
review 1 thru 7.

Section 2: Chapter
review 12, 13. Section

3: Chapter review 19
thru 25.

Chapter Fourteen [Acids and Bases] - Wattsburg

Chapter 14 (Acids and
Bases) - Part 1 -

YouTube. Major topics:

Arrhenius vs. Bronsted-

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Chapter 14 Acids And Bases

Lowry definition of acids and bases, conjugate acid/base, acid dissociation constant (K_a), & strong vs weak...

Chapter 14 (Acids and Bases) - Part 1

Acids and Bases:

Chapter 14 & 15 . HW:

- Read Ch 14:
- Fill in as much of the acid base table as you can, as you read . Acid base conductivity and reactivity

Conducvity,

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Chapter 14 Acids And Bases

Reac'tivity,,

Hydrochloric*acid* high

high Acidic*acid* **

low* medium

Dis4lled*water* none*

none*

Ammoniumhydroxide

med* none*

Sodiumhydroxide* high

none* *

Acids and Bases:

Chapter 14 & 15

Chapter 14: Acids and

Bases 1. Acids and

Bases. 1. Which of the

following is the

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equilibrium constant expression for the dissociation of the weak acid HOCl?
e) none of these. 2. In which of the following reactions does the H_2PO_4 ion act as an acid?
d) The ion cannot act as an acid.

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ecf8427e.

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